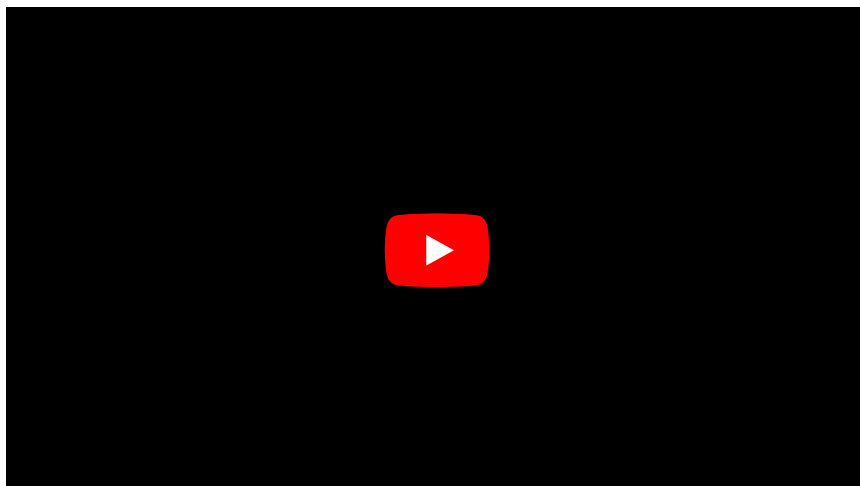


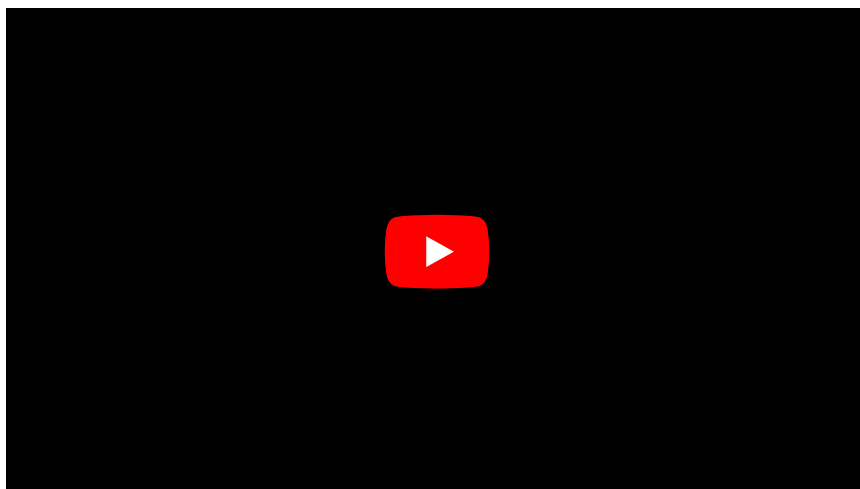
# Osmolarity

**Osmolarity** is the amount of all osmotically active substances in 1 L of solution. It is equal to the product of molarity and the number of particles formed by the dissociation of one molecule, expressed in  $\text{osm/L}$ .

Unlike osmolality, osmolarity is the amount of substances in the total volume of the solution (including dissolved substances). For strongly diluted water solutions, the value of osmolarity and osmolality hardly differ, since 1 L of water is approximately equal to 1 kg of water and the amount of dissolved substances can be neglected.



## Osmolarity:



## Links

### Related articles

- Osmolality
- Osmotic pressure
- molarity

### Bibliography

- ŠVÍGLEROVÁ, Jitka. *Osmolarita* [online]. The last revision 18. 2. 2009, [cit. 10.11.2010]. <<https://web.archive.org/web/20160416225455/http://wiki.lfp-studium.cz/index.php/Osmolarita>>.

